

Chapter 7 Ionic And Metallic Bonding Vocabulary Review Answers

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Chapter 7 "Ionic and Metallic Bonding" Pre-AP Chemistry Charles Page High School
Stephen L. Cotton. Section 7.1 - Ions OBJECTIVES: - Determine the number of
valence electrons in an atom of a representative element. Section 7.1 - Ions
OBJECTIVES: - Explain how the octet rule applies to atoms of metallic and
nonmetallic elements.

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SECTION 7.1 IONS (pages 187-193)

Chapter 7 Ionic and Metallic Bonding59 SECTION 7.1 IONS (pages 187-193) This
section explains how to use the periodic table to infer the number of valence
electrons in an atom and draw its electron dot structure. It also describes the
formation of cations from metals and anions from nonmetals.

Chapter 7 Ionic And Metallic Bonding Answer Key

Lesson 7.1 Reading and Study Workbook • Copyright © Pearson Education, Inc., or
its affiliates. All Rights Reserved. 83. Ionic and Metallic Bonding. BONDING AND
INTERACTIONS. 7.1 Ions. EssentialUnderstanding Ions form when atoms gain or
lose valence electrons, becoming electrically charged. Lesson Summary. Valence
ElectronsValence electrons are the electrons in the outermost occupied energy
level and are involved in ion formation. For a representative element, the group
number equals the

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Chapter 7 Ionic and Metallic Bonding 163 SECTION 7.2 IONIC BONDS AND IONIC COMPOUNDS 1. Use electron dot structures to predict the formula of the ionic compounds formed when the following elements combine. a. sodium and bromine d. aluminum and oxygen b. sodium and sulfur e. barium and chlorine c. calcium and iodine 2. Which of these combinations of elements are most likely to react to form ionic

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Chapter 7 - Ionic and Metallic Bonding Chapter 7 - Ionic and Metallic Bonding - 7.1 Ions - 7.1 Lesson Check - Page 199: 1 Answer To find the number of valence electrons in an a representative element simply look at the group number. Chemistry (12th Edition) Chapter 7 - Ionic and Metallic

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halide ion ionic bonds ionic compounds metallic bonds octet rule valence e Section 7 1 Ions OBJECTIVES . Determine the number of valence electrons in an atom of a representative Section 7 1 Ions OBJECTIVES . Explain how the octet rule applies to atoms of metallic and nonmetallic Section 7 1 Ions OBJECTIVES .

Ionic And Metallic Bonding Chapter 7 Worksheet Answers

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BONDING AND INTERACTIONS

Chapter 7 Review “Ionic and Metallic Bonding” Chapter 7 Review How many electrons does nitrogen gain in order to achieve a noble-gas electron configuration? How does oxygen obey the octet rule when reacting to form compounds? The electron configuration of a fluoride ion (F¹⁻) is _____. Chapter 7 Review How many valence electrons are in an

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Section 7.3 – Bonding in Metals. The valence electrons of metal atoms can be modeled as a sea of electrons. Metallic bonds consist of the attraction of the free-floating valence electrons for the positively charged metal ions. Metals are good conductors and malleable because of their mobile electrons.

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Chapter 7 - Ionic and Metallic Bonding - 7 Assessment - Page 215: 49 Answer 3 crystal arrangements are body-centered cubic (like sodium), face-centered cubic (like copper), and hexagonal close-packed (like zinc).

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Chapter 7 - Ionic and Metallic Bonding

Section 7.2 – Ionic Bonds and Ionic Compounds. Compounds composed of cations

and anions are called ionic compounds. Ionic compounds are usually composed of a metal and a nonmetal. In contrast, molecular compounds are generally composed of nonmetals only. Although they are composed of ions, ionic compounds are electrically neutral.

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