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knowledge of classical physics, the model's success lay in explaining the Rydberg formula for the spectral emission lines of atomic hydrogen. The Bohr Model | Introduction to Chemistry The Bohr model gives almost exact results only for a system where two charged points orbit each other at speeds much less than that of light. This not only involves one-electron systems such as the hydrogen atom, singly ionized helium, and doubly ionized lithium, but it includes positronium and Rydberg states of any atom where one electron is far away from everything else. Bohr model -

Wikipedia Bohr model, description of the structure of atoms, especially that of hydrogen, proposed (1913) by the Danish physicist Niels Bohr. The Bohr model of the atom, a radical departure from earlier, classical descriptions, was the first that incorporated quantum theory and was the predecessor of wholly quantum-mechanical models. Bohr model | Description & Development | Britannica Gizmo Warm-up In 1913, Niels Bohr proposed that the unique spectral lines created by different elements were related to the way electrons were arranged around the nucleus. The Bohr Model: Introduction Gizmo™ explores this connection. The laser shown in the Gizmo can emit photons, or particles of light, at a variety of wavelengths. Student Exploration Bohr Model- Introduction (ANSWER KEY ... Bohr Model of Hydrogen The simplest example of the Bohr Model is for the hydrogen atom ($Z = 1$) or for a hydrogen-like ion ($Z > 1$), in which a negatively charged electron orbits a small positively charged nucleus. Electromagnetic energy will be absorbed or emitted if an electron moves from one orbit to another. Bohr Model of the Atom - Overview and

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